

(19)



Europäisches Patentamt

European Patent Office

Office européen des brevets



(11)

EP 1 222 954 A1

(12)

EUROPEAN PATENT APPLICATION

(43) Date of publication:
17.07.2002 Bulletin 2002/29

(51) Int Cl.7: B01D 61/48, C02F 1/00

(21) Application number: 01310901.2

(22) Date of filing: 27.12.2001

(84) Designated Contracting States:
AT BE CH CY DE DK ES FI FR GB GR IE IT LI LU
MC NL PT SE TR
Designated Extension States:
AL LT LV MK RO SI

(72) Inventors:
• Sato, Shin, c/o Kurita Water Industries Ltd.
Shinjuku-ku, Tokyo (JP)
• Moribe, Takayuki,
c/o Kurita Water Industries Ltd.
Shinjuku-ku, Tokyo (JP)

(30) Priority: 05.01.2001 JP 2001000678

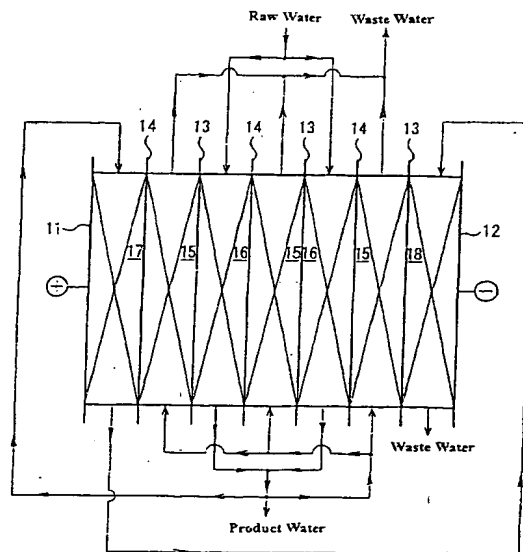
(71) Applicant: KURITA WATER INDUSTRIES LTD.
Shinjuku-ku Tokyo (JP)

(74) Representative: Jenkins, Peter David et al
PAGE WHITE & FARRER
54 Doughty Street
London WC1N 2LS (GB)

(54) Method and apparatus for electrodeionization of water

(57) An electrodeionization apparatus has, an anolyte compartment 17 having an anode 11, a catholyte compartment 18 having a cathode 12, concentrating compartments 15, and desalting compartments 16. The concentrating compartments 15 and the desalting compartments 16 are alternately formed between the anolyte compartment 17 and the catholyte compartment 18 by alternately arranging a plurality of anion-exchange membranes 13 and a plurality of cation-exchange membranes 14. The desalting compartments 16 are filled with ion-exchanger and the concentrating compartments 15 are filled with ion-exchanger, activated carbon, or electric conductor. Electrode water flows into the anolyte compartment 17 and the catholyte compartment 18. Concentrated water is introduced into the concentrating compartments 15. Raw water is fed into the desalting compartment 16 to produce the deionized water from the desalting compartment 16. Water containing silica or boron at a lower concentration than the raw water is introduced into the concentrating compartments 15 as the concentrated water in a direction from a side near an outlet for the deionized water toward a side near an inlet for the raw water of the desalting compartments 16. At least a part of concentrated water flowing out of the concentrating compartments 15 is discharged out of a circulatory system.

Fig.1



Description

[0001] The present invention relates to an electrodeionization apparatus, a method of operating an electrodeionization apparatus, and a system for producing ultra pure water.

5 [0002] Deionized water is used for various purposes, for example, in plants such as for semiconductor production and liquid crystal display production, in industrial facilities such as for pharmaceutical industry, food industry, and electric power industry, even in households, and in laboratories. Electrodeionization apparatuses are frequently used to produce deionized water as described in Japanese Patent No. 1782943, Japanese Patent No. 2751090, and Japanese Patent No. 2699256. A conventional electrodeionization apparatus of Fig.2 includes electrodes which consist of an
10 anode 11' and a cathode 12', anion-exchange membranes 13 and cation-exchange membranes 14'. The membranes are alternately arranged in such a manner as to alternately form concentrating compartments 15' and desalting compartments 16' between the anode and the cathode. The desalting compartments 16' are filled with anion-exchanger and cation-exchanger made of ion exchange resin, ion exchange fibers, or graft exchanger. In the desalting compartments 16', the anion-exchanger and cation-exchanger are in the mixed state or multiple-layered state.

15 [0003] Ions flowing into the desalting compartments 16' react with the ion exchanger according to the affinity, concentration, and mobility of the ions and move through the ion exchanger in a direction of potential gradient. The ions further pass through the membranes to hold neutralization of charges in all of the compartments. The ions decrease in the desalting compartments 16' and increase in the concentrating compartments 15' because of the semi-permeability of the membranes and the polarities of potential gradient. This means that cations permeate the cation-exchange
20 membranes 14' and anions permeate the anion-exchange membranes 13' so that the cations and anions are concentrated in the concentrating compartments 15'. Therefore, deionized water (pure water) as product water is recovered from the desalting compartments 16'.

[0004] Electrode water flows through an anolyte compartment 17' and a catholyte compartment 18'. The water flowing out of the concentrating compartments 15' (concentrated water) and having high ion concentration is as the electrode
25 water in order to ensure the electric conductivity.

[0005] Raw water is introduced into the desalting compartments 16' and the concentrating compartments 15'. Deionized water (pure water) is taken out from the desalting compartments 16'. Concentrated water in which ions are concentrated is discharged from the concentrating compartments 15'. A part of the concentrated water is circulated into the inlets of the concentrating compartments 15' by a pump (not shown) in order to improve the product water recovery.
30 Another part of the concentrated water is supplied to the inlet of the anolyte compartment 17'. The reminder of the concentrated water is discharged as waste water out of a circulatory system in order to prevent the ion concentration in the circulatory system. Water flowing out of the anolyte compartment 17' is supplied to the inlet of the catholyte compartment 18'. Water flowing out of the catholyte compartment 18' is discharged as waste water out of the circulating system.

35 [0006] The pH in the anolyte compartment 17' is lowered due to H^+ generated by dissociation of water. On the other hand, the pH in the catholyte compartment 18' is increased due to generation of OH^- . The acid water flowing out of the anolyte compartment 17' flows into the catholyte compartment 18' so that alkalinity in the catholyte compartment 18' can be neutralized, thereby eliminating damages due to scale formation.

[0007] Filling activated carbon or ion-exchange resin into electrode compartments is disclosed in USP 5,868,915.

40 [0008] The above conventional electrodeionization apparatus do not remove silica and boron at extremely high ratio.

[0009] It is an aim of the present invention to provide an electrodeionization apparatus which removes silica and boron at extremely high ratio, a method of operating the same, and a system employing the electrodeionization apparatus for producing ultra pure water.

[0010] A method for electrodeionization according to a first aspect of the present invention employs an electrodeionization apparatus which has an anolyte compartment having an anode, a catholyte compartment having a cathode, at least one concentrating compartment, and at least one desalting compartment. The concentrating compartment(s) and the desalting compartment(s) are formed between the anolyte compartment and the catholyte compartment by alternately arranging at least one anion-exchange membrane(s) and at least one cation-exchange membrane(s). The desalting compartment(s) is (are) filled with ion-exchanger, and the concentrating compartment(s) is (are) filled with ion-exchanger, activated carbon, or an electric conductor. Electrode water flows into the anolyte compartment(s) and the catholyte compartment(s). Concentrated water flows into the concentrating compartment(s). Raw water flows into the desalting compartment(s) and the deionized water flows out of the desalting compartment(s). The concentrated water includes silica or boron at a lower concentration than the raw water. The concentrated water flows into the concentrating compartment(s) at a side near an outlet for the deionized water of the desalting compartment and flows out from the concentrating compartment(s) at a side near an inlet for the raw water of the desalting compartment. At least a part of the concentrated water flowing out of the concentrating compartments is discharged out of the circulating system.
55

[0011] An electrodeionization apparatus according to a second aspect of the present invention has an anolyte compartment having an anode, a catholyte compartment having a cathode, at least one concentrating compartment, and

at least one desalting compartment. The concentrating compartment(s) and the desalting compartment(s) are alternately formed between the anolyte compartment and the catholyte compartment by alternately arranging at least one anion-exchange membrane(s) and at least one cation-exchange membrane(s). The desalting compartment(s) is (are) filled with ion-exchanger, and the concentrating compartment(s) is (are) filled with ion-exchanger, activated carbon, or an electric conductor. The electrodeionization apparatus further has a device for introducing electrode water into the anolyte compartment and the catholyte compartment; a concentrated water introducing device for introducing concentrated water into the concentrating compartment(s); and a device for introducing raw water into the desalting compartment(s) to produce the deionized water. The concentrated water introducing device introduces water containing silica or boron at a lower concentration than the raw water into the concentrating compartment(s) at a side near an outlet for the deionized water of the desalting compartment(s). The concentrated water flows out of the concentrating compartments at a side near an inlet for the raw water of the desalting compartment(s). At least a part of concentrated water flowing out of the concentrating compartments is discharged out of a circulating system.

[0012] A system for producing ultra pure water of the present invention has the above electrodeionization apparatus of the second aspect of the present invention.

[0013] Decreasing in silica or boron concentration in the concentrated water flowing into the concentrating compartment near the outlet for the product water leads to decrease in silica or boron concentration in the product water.

[0014] By introducing concentrated water containing silica or boron at a lower concentration than the raw water into the desalting compartments at a side near the outlet for the deionized water (product water) in a direction toward a side near the inlet for the raw water, the silica or boron concentration of product water is significantly decreased.

[0015] By introducing water having low electric conductivity and high resistivity in the concentrating compartments, the electric resistance of the water in the concentrating compartments is increased.

[0016] The ion exchanger such as ion-exchange resin, activated carbon, or electric conductor filled in the concentrating compartments reduces the electric resistance of the concentrating compartments, allowing much electric current to flow.

[0017] H^+ ions and OH^- ions produced by dissociation of water among the ion-exchange resins carry electric charges, so that voltage between electrodes is prevented from rising so as to allow enough current to flow between the electrodes even when water having high resistivity such as ultra pure water flows in the electrodeionization apparatus.

[0018] Embodiments of the present invention will now be described by way of example only, with reference to the accompanying drawings, in which:-

Fig. 1 is a schematic sectional view showing an electrodeionization apparatus according to an embodiment of the present invention;

Fig. 2 is a schematic sectional view showing a conventional electrodeionization apparatus;

Fig. 3a is a perspective view schematically showing an electrodeionization apparatus according to another embodiment of the present invention;

Fig. 3b is a schematic flow diagram of the apparatus of Fig. 3a; and

Fig. 4 is a schematic flow diagram of a system for producing ultra pure water in which the electrodeionization apparatus of the present invention is employed.

[0019] An electrodeionization apparatus shown in Fig. 1 has a plurality of anion-exchange membranes 13 and a plurality of cation-exchange membranes 14 which are alternately arranged between the electrodes (anode 11, cathode 12), concentrating compartments 15, and desalting compartments 16. The concentrating compartments 15 and the desalting compartments 16 are each defined between the membranes 13 and 14 and are therefore alternately arranged between the electrodes. The desalting compartments 16 are filled with anion-exchanger and cation-exchanger made of ion exchange resin, ion exchange fibers, or graft exchanger. In the desalting compartments 16, the anion-exchanger and cation-exchanger are filled in the mixed state or multiple-layered state.

[0020] The concentrating compartments 15, anolyte compartment 17, and catholyte compartment 18 are filled with electric conductive media such as ion exchanger, activated carbon, or metal.

[0021] Raw water is introduced into the desalting compartments 16. Product water is taken out from the desalting compartments 16. A part of the product water flows into the concentrating compartments 15 in a direction opposite to the flowing direction of the desalting compartments 16 i.e. in single-pass counter-flow manner. Water flowing out of the concentrating compartments 15 is discharged out of a system of the apparatus. The concentrating compartments 15 are provided with inlets on the same side of the outlets for the product water of the desalting compartments 16 and provided with outlets on the same side of the inlets for the raw water of the desalting compartments 16. Another part of the product water is supplied to the inlet of the anolyte compartment 17. Water flowing out of the anolyte compartment 17 is supplied to the inlet of the catholyte compartment 18. Water flowing out of the catholyte compartment 18 is discharged as waste water out of the system.

[0022] By introducing product water into the concentrating compartments 15 in the single-pass counter-flow manner

relative to the desalting compartments 16, the concentrated water in the concentrating compartment 15 near the outlets for product water has the lowest ion concentration, whereby the ion diffusion to the desalting compartments 16 due to the concentration diffusion is restricted, and the ions are removed at a high rate. Especially, silica and boron ions are removed at an extremely high rate.

5 [0023] In the conventional electrodeionization apparatus, a part of concentrated water (water flowing out of the concentrating compartments) is discharged as shown in Fig. 2 in order to improve the product water recovery. The remainder of the concentrated water is circulated to the inlets of the concentrating compartments. In the conventional apparatus, the LV (line velocity) in the concentrating compartments is 80 m/hr or more.

10 [0024] Since the ion exchanger is filled in the concentrating compartments in the present embodiment of the invention, deionizing property can be ensured even when the LV in the concentrating compartments is 20 m/hr or less. The reasons will be described as follows. When a spacer is placed in each concentrating compartment, it is required to disperse condensed silica and condensed boron on membranes by water flows in the concentrating compartment. However, when the ion exchanger is filled in the concentrating compartments, ions are dispersed through the ion exchanger, thereby eliminating the necessity of high line velocity (LV).

15 [0025] Because the high line velocity is not necessary, the product water recovery can be higher than the conventional one even though the concentrated water flows in the single-pass manner. In addition, no circulation pump is required, whereby the apparatus can be more economical.

[0026] The silica concentration of concentrated water flowing at locations corresponding to the outlets for product water is preferable to as low as possible. It is desirable that the silica concentration is not greater than 1000 times as great as that of the product water.

20 [0027] Though the filler for the concentrating compartments may be activated carbon to ensure required current, ion exchanger is preferably used rather than activated carbon because of the ion diffusion action as mentioned above.

[0028] In the electrodeionization apparatus shown in Fig. 1, a part of the product water is also supplied to the electrode compartments 17, 18. In order to ensure desired current, the electrode compartments 17, 18 are filled with ion exchanger, activated carbon, or metal as electric conductive media in the same manner as the concentrating compartments 15. This makes the applied voltage constant regardless of water quality. Therefore, even when high quality water such as ultra pure water is flowed, desired current can be ensured.

25 [0029] In the electrode compartments, particularly in the anolyte compartment, oxidizing agent such as chlorine and ozone are produced. Therefore, the filler is more preferably activated carbons for long-term use rather than ion-exchange resin. As shown in Fig. 1, it is preferable to feed product water having little or no Cl⁻ into the electrode compartments in view of long-term stabilization of the filler and the electrodes because production of chlorine in the electrode compartments is prevented.

30 [0030] The electrode compartments may not be filled with the aforementioned filler. For example, each electrode plate is provided, on a water flowing side, with a structure being porous so that electrode water permeates the porous portion. In this case, the electrode plates and the electrode compartments can be formed as an integral part, thereby facilitating the assembly.

35 [0031] When the concentrated water is circulated in the concentrating compartment the inside of which is not divided, the concentration of silica and boron should be higher near the outlets for product water. When concentrated water is circulated in a concentrating compartment the inside of which is divided as shown in Figs. 3a, 3b, the concentration at the side near the outlet for the concentrating compartment is higher than that at the side near the outlet for the concentrating compartment so that the quality of product water is substantially equal to the quality of product water obtained in the single-pass counter-flow arrangement of Fig. 1.

40 [0032] Electrodeionization apparatuses shown in Figs. 3a, 3b have an anode 11 and a cathode 12. Cation-exchange membranes and anion-exchange membranes are alternately arranged between the anode 11 and the cathode 12 to define alternately a concentrating compartment 15 and desalting compartments 16. The concentrating compartment 15 is divided into two or more (two in Figs. 3a, 3b) concentrated water flowing sections 15A, 15B by a partition 15S. The flowing direction of concentrated water in each concentrated water flowing section 15A, 15B crosses the flowing direction in the desalting compartments 16.

45 [0033] Each desalting compartment 16 has an inlet at the top and an outlet at the bottom in Fig. 3a so that water flows downwardly in a vertical direction.

50 [0034] The concentrating compartment 15 is provided with the partition 15S extending in a direction crossing the flowing direction in the desalting compartments 16. Although the direction is perpendicular to the flowing direction of the desalting compartments 16 in Fig. 3a, the term "perpendicular" includes a range of angle between 80°-100°. The inside of the concentrating compartment 15 is divided into two stages which are arranged vertically in Fig. 3a, by the partition 15S. Water flows from the front to the back in Fig. 3a in the respective concentrated water flowing sections 15A, 15B.

55 [0035] As shown in Fig. 3b, a part of the product water flowed out of the desalting compartments is introduced into a circulatory system of the concentrated water flowing section 15B in which the circulation is conducted by a pump.

The part of product water is thus circulated in the concentrated water flowing section 15B near the outlets for product water. A part of circulating concentrated water from the circulatory system is introduced into a circulatory system of the concentrated water flowing section 15A in which the circulation is conducted by a pump. The part of circulating concentrated water is thus circulated in the concentrated water flowing section 15A near the inlets for raw water. A part of circulating concentrated water from the concentrated water flowing section 15A near the inlets for raw water is discharged out of the circulatory system.

[0036] In the electrodeionization apparatus of Figs. 3a, 3b, after a part of product water enters into a circulatory system of the concentrated water flowing section 15B near the outlet for product water and is circulated therein, a part of circulated water from the concentrated water flowing section 15B enters into a circulatory system of the concentrated water flowing section 15A near the inlet for raw water, is circulated therein, and is discharged out of the circulatory system. This means that concentrated water is flowed from the side of the outlets for product water to the side of the inlets for raw water and, after that, is partially discharged out of the circulatory system. Accordingly, the apparatus exhibits the same effects as the case shown in Fig. 1 in which water in the concentrating compartment flows in the single-pass counter-flow manner relative to the desalting compartment.

[0037] There may be three or more concentrated water flowing sections defined by partitions in the concentrating compartment. In view of the increase in number of partitions or parts and complexity of the apparatus structure, the concentrating compartment is preferably divided into two or three concentrated water flowing sections.

[0038] The smallest possible thickness of the desalting compartment is preferable for removing not only silica but also boron in the electrodeionization apparatus. The thickness of the desalting compartment is preferably 5 mm or less. However, in view of water permeability and ease of manufacturing, the thickness is preferably 2 mm or more in practice.

[0039] According to the present invention, required current passes between the electrodes and silica and boron are removed at high rate while eliminating the influence of concentration diffusion. The required current passes in the apparatus having the above-described concentrating compartment and the electrode compartments. The current required for increasing the removal ratio of silica and boron is a current value corresponding to current efficiency of 10% or less. To obtain the removal ratio of silica and boron of 99.9% or more, the required current is a current value preferably corresponding to current efficiency of 5% or less. The current efficiency is expressed by the following equation:

$$\text{Current Efficiency (\%)} = 1.31 [\text{flow rate per cell (L/min)}]$$

$$[\text{equivalent conductivity of raw water } (\mu\text{S/cm})] - [\text{equivalent conductivity of treated water } (\mu\text{S/cm})] / \text{current (A)}$$

[0040] According to the electrodeionization apparatus of the present invention, required current can be ensured even when water having high resistivity is fed as raw water into the electrodeionization apparatus and it is required to decrease further only silica and boron in the raw water.

[0041] It should be noted that if no current flows in any one of concentrating compartments and electrode compartments in a conventional electrodeionization apparatus, current does not flow through the apparatus.

[0042] On the other hand, the apparatus of the present invention can remove silica and boron from raw water having high resistivity. Therefore, the electrodeionization apparatus of the present invention can treat various kinds of water.

[0043] For example, the electrodeionization apparatus can be employed as a primary pure water producing apparatus in a semiconductor plant. Even when product water produced by the primary pure water producing apparatus is consumed in small quantities and the remainder is returned to be circulated as raw water so as to make the raw water have high resistivity, required current can be ensured. Therefore, the apparatus can be stably started up.

[0044] Even when a plurality of electrodeionization apparatuses of the present invention are arranged in series and raw water is introduced in these apparatus, required current for the subsequent apparatus can also be secured.

[0045] The electrodeionization apparatus of the present invention can be employed also as a secondary pure water producing system called sometimes "sub-system" in an ultra pure water producing process. Even when water having resistivity of 10 MΩ·cm or more is fed as raw water into this apparatus, required current can be ensured. Therefore, the electrodeionization apparatus of the present invention can be employed as an alternative to a demineralizer (non-regenerative mixed-bed ion exchange apparatus) as shown in Fig. 4.

[0046] In the sub-system of Fig. 4, primary pure water having resistivity of 10 MΩ·cm or more is introduced into a desalting compartment 2A of the electrodeionization apparatus 2 via an ultraviolet oxidizing apparatus 1. Product water from the electrodeionization apparatus 2 is treated at an ultrafiltration membrane separation apparatus 3 so as to produce ultra pure water. Taking the entire water balance into consideration, concentrated water from the ultrafiltration membrane separation apparatus 3 may be employed as feed water to the concentrating compartment 2B of the elec-

trodeionization apparatus 2. Therefore, efficiency for water utilization of a system as a whole can be increased.
 [0047] Hereinafter, examples and comparative examples will be described.

Example 1

[0048] City water was first filtered with activated carbon, treated by reverse osmosis separator, and further treated with membrane deaerator to become raw water having quality shown in Table 1. The raw water was deionized by the electrodeionization apparatus of Fig. 1 having two desalting compartments and three concentrating compartments and having the following specification. The electrodeionization apparatus was operated under the following operating conditions.

[Specification of Electrodeionization Apparatus]

[0049]

Height of Desalting Compartment : 66 cm
 Thickness of Desalting Compartment: 2.5 mm
 Thickness of Concentrating Compartment: 2.5 mm
 Filler for Desalting Compartment: Mixed ion-exchange resin of anion-exchange resin : cation-exchange resin = 7 : 3 (by volume)
 Filler for Concentrating Compartment : Mixed ion-exchange resin of anion-exchange resin : cation-exchange resin = 7 : 3 (by volume)
 Filler for Anolyte and Catholyte Compartments : Activated carbon

[Operating Conditions]

[0050]

Current : 2A (Current Efficiency 4%)
 Desalting Compartment SV : 130 hr⁻¹
 Concentrating Compartment LV : 13 m/hr
 Product Water: 60 L/hr
 Water in Concentrating Compartment: 9 L/hr
 Water in Electrode compartment: 5 L/hr
 Water Recovery: 81% (= 60 ÷ (60 + 9 + 5) × 100)

[0051] Product water flowed at 74L/hr, and 9 L/hr of the product water was introduced into the concentrating compartments in the single-pass counter-flow manner and 5 L/hr was supplied to the anolyte compartment and then flowed in the catholyte compartment.

[0052] Resultant product water was high-purity water in which both silica and boron were removed to a high degree so that both residual silica and boron were under the detection limit as shown in Table 1.

[0053] Applied voltage for the above experiment was 9.7V in total, that is, desalting compartment: 1.28V × 2, concentrating compartment: 1.28V × 3, anolyte compartment: 1.4V, and catholyte compartment: 1.9V.

Table 1

	Raw Water	Product Water
Silica (ppb)	300	0.1 or less (under detection limit)
Boron (ppb)	10	0.1 or less (under detection limit)
Conductivity (μS/cm)	10	-
Resistivity (MΩ·cm)	-	18

Example 2

[0054] Deionization was conducted under the same conditions as Example 1 except that the raw water was prepared by adding sodium silicate into ultra pure water to have sodium silicate concentration of 300 ppb (SiO₂ conversion).

[0055] The raw water contains low ion concentration, so that it is difficult to ensure required current. However, the product water had high-purity such that silica concentration was 0.1 ppb or less (under the detection limit).

[0056] Applied voltage for the above experiment was also 9.7 volts in total, that is, desalting compartment: $1.28V \times 2$, concentrating compartment: $1.28V \times 3$, anolyte compartment: 1.4V, and catholyte compartment: 1.9V.

Comparative Example 1

[0057] An experiment was operated by using the apparatus as shown in Fig. 2 under the same conditions as Example 1 except that the concentrated water rate was 60 L/hr and the concentrated water discharge rate was 9 L/hr.

[0058] The product water had resistivity of 17 M Ω ·cm and silica concentration of 6 ppb.

[0059] As described above, according to the present invention, the electrodeionization apparatus produces high-purity product water in which both silica and boron are removed to a high degree. Conventional electrodeionization apparatuses could not sufficiently remove silica and boron.

Claims

1. A method of operating an electrodeionization apparatus which includes an anolyte compartment having an anode, a catholyte compartment having a cathode, at least one concentrating compartment, and at least one desalting compartment,
the concentrating compartment and the desalting compartment being formed between the anolyte compartment and the catholyte compartment by arranging at least one anion-exchange membrane and at least one cation-exchange membrane,
the desalting compartment being filled with ion-exchanger,
the concentrating compartment being filled with at least one of ion-exchanger, activated carbon, and electric conductor;
wherein the electrodeionization apparatus produces deionized water by supplying electrode water into the anolyte compartment and the catholyte compartment; supplying concentrated water into the concentrating compartment; and feeding raw water into the desalting compartments to produce the deionized water from the desalting compartment;
wherein the concentrated water includes at least one of silica and boron at a lower concentration than the raw water, and the concentrated water is introduced into the concentrating compartment at a side near an outlet for deionized water of the desalting compartment and flows out of the concentrating compartments at a side near an inlet for the raw water of the desalting compartment, and
wherein at least a part of the concentrated water flowing out of the concentrating compartment is discharged out of a circulatory system.
2. A method of operating an electrodeionization apparatus as claimed in claim 1, wherein the concentrated water flows in the concentrating compartment in single-pass counter-flow manner relative to the water in the desalting compartment.
3. A method of operating an electrodeionization apparatus as claimed in claim 1, wherein at least one of desalted water of the electrodeionization apparatus, treated water produced by further treating the desalted water by another apparatus such as an ion exchange apparatus, and ultra pure water is introduced into the concentrating compartment as the concentrated water.
4. A method of operating an electrodeionization apparatus as claimed in claim 1, wherein the anolyte compartment and the catholyte compartment are filled with at least one of activated carbon, ion-exchanger, and electric conductor.
5. A method of operating an electrodeionization apparatus as claimed in claim 1, wherein
the anode is in contact with a cation-exchange membrane which defines the anolyte compartment,
the cathode is in contact with an anion-exchange membrane which defines the catholyte compartment, and
the anode and the cathode are each provided, at least at a side being in contact with the corresponding membrane, with a porous structure having continuous multiple apertures through which electrode water flows.
6. A method of operating an electrodeionization apparatus as claimed in claim 1, wherein the line velocity (LV) of the concentrated water in the concentrating compartment is 20 m/hr or less.

7. A method of operating an electrodeionization apparatus as claimed in claim 1, wherein the thickness of the desalting compartment is 2-5 mm.

5 8. A method of operating an electrodeionization apparatus as claimed in claim 1, wherein the current has a current value that the current efficiency of the electrodeionization apparatus expressed by the following equation is 10% or less:

$$\text{Current Efficiency (\%)} = 1.31 [\text{flow rate per cell (L/min)}]$$

10

$$\frac{[\text{equivalent conductivity of raw water } (\mu\text{S/cm})] - [\text{equivalent conductivity of treated water } (\mu\text{S/cm})]}{\text{current (A)}}$$

15 9. A method of operating an electrodeionization apparatus as claimed in claim 8, wherein the current has a current value that the current efficiency is 5% or less.

10. An electrodeionization apparatus comprising:

20 an anolyte compartment having an anode;
a catholyte compartment having a cathode;
at least one concentrating compartment;
at least one desalting compartment wherein the concentrating compartments and the desalting compartments
are formed between the anolyte compartment and the catholyte compartment by arranging at least one anion-
25 exchange membrane and at least one cation-exchange membrane;
ion-exchanger with which the desalting compartment is filled;
at least one of ion-exchanger, activated carbon, and electric conductor which fills the concentrating compartment;
a device for introducing electrode water into the anolyte compartment and the catholyte compartment, respectively;
30 a concentrated water introducing device for introducing concentrated water into the concentrating compartments; and
a device for feeding raw water into the desalting compartment to produce the deionized water,

35 wherein the concentrated water introducing device introduces water containing at least one of silica and boron at a lower concentration than the raw water into the concentrating compartments at a side near an outlet for the deionized water of the desalting compartment;

the concentrated water introducing device makes the concentrated water flow out of the concentrating compartment at a side near an inlet for the raw water of the desalting compartment; and

40 the concentrated water introducing device discharges at least a part of the concentrated water flowing out of the concentrating compartments out of a circulatory system.

11. An electrodeionization apparatus as claimed in claim 10, wherein the concentrated water introducing device introduces the concentrated water into the concentrating compartment in single-pass counter-flow manner relative to the water in the desalting compartment.

12. An electrodeionization apparatus as claimed in claim 10, wherein one of the desalted water produced by the electrodeionization apparatus, treated water prepared by further treating the desalted water by another apparatus such as an ion exchange apparatus, and ultra pure water is introduced into the concentrating compartment as the concentrated water.

13. An electrodeionization apparatus as claimed in claim 10, wherein the anolyte compartment and catholyte compartment are filled with at least one of activated carbon, ion-exchanger, and electric conductor.

55 14. An electrodeionization apparatus as claimed in claim 10, wherein the anode is in contact with a cation-exchange membrane which defines the anolyte compartment,
the cathode is in contact with an anion-exchange membrane which defines the catholyte compartment, and the anode and the cathode are each provided, at least at a side being in contact with the corresponding

membrane, with a porous structure having continuous multiple apertures through which electrode water flows in the anolyte compartment and the catholyte compartment.

5 15. An electrodeionization apparatus as claimed in claim 10, wherein the line velocity (LV) of the concentrated water in the concentrating compartment is 20 m/hr or less.

16. An electrodeionization apparatus as claimed in claim 10, wherein the thickness of the desalting compartment is 2-5 mm.

10 17. An electrodeionization apparatus as claimed in claim 10, wherein the current has a current value that the current efficiency of the electrodeionization apparatus expressed by the following equation is 10% or less:

$$\text{Current Efficiency (\%)} = 1.31 [\text{flow rate per cell (L/min)}]$$

15

$$[[\text{equivalent conductivity of raw water } (\mu\text{S/cm})] - [\text{equivalent}$$

$$\text{conductivity of treated water } (\mu\text{S/cm})]] / \text{current (A)}$$

20 18. An electrodeionization apparatus as claimed in claim 17, wherein the current has a current value that the current efficiency is 5% or less.

19. A system for producing ultra pure water comprising the electrodeionization apparatus as claimed in claim 10.

25 20. A system for producing ultra pure water as claimed in claim 19, further comprising an ultrafiltration membrane separation apparatus into which deionized water from the electrodeionization apparatus is introduced, wherein concentrated water from the ultrafiltration membrane separation apparatus is introduced into the concentrating compartment of the electrodeionization apparatus.

30

35

40

45

50

55

Fig.1

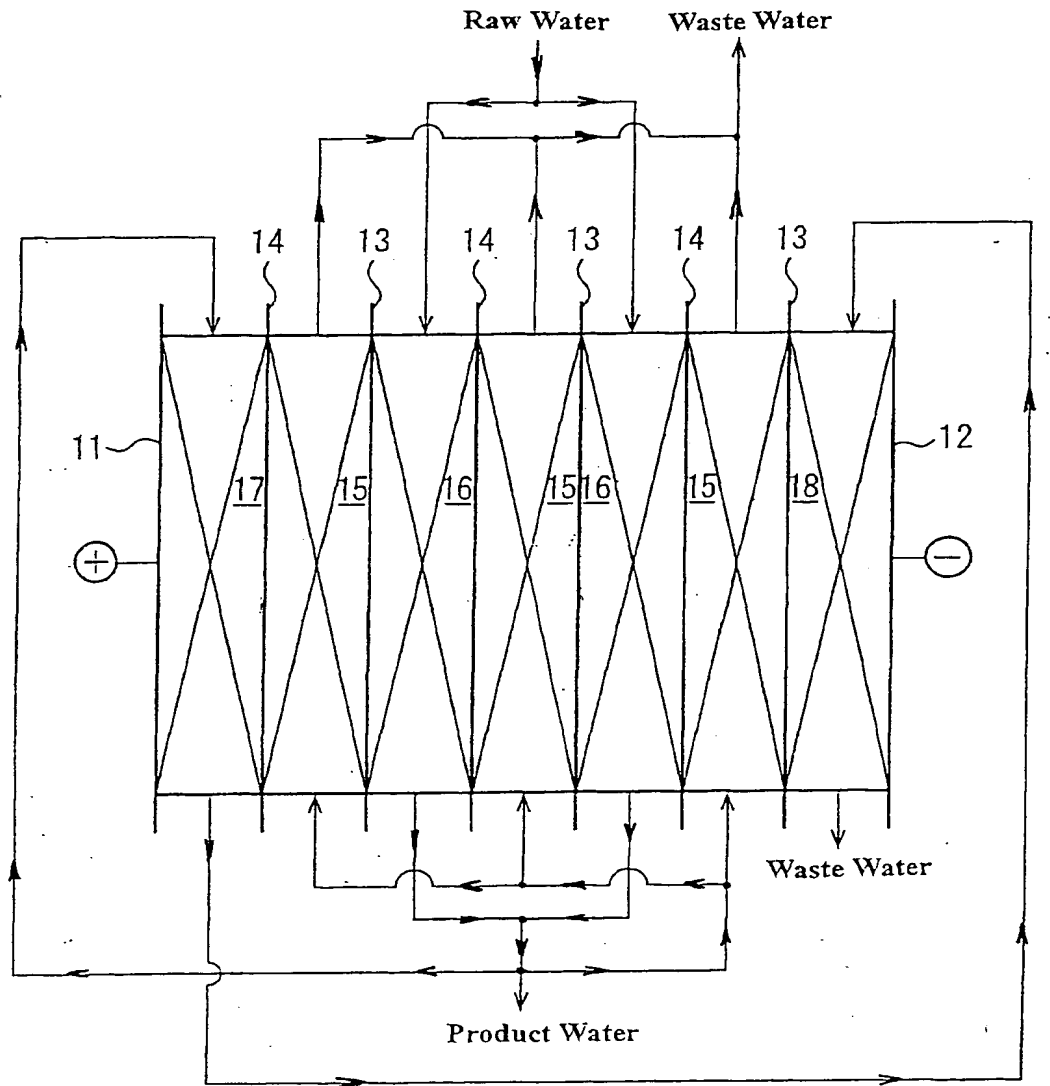


Fig.2 (Prior Art)

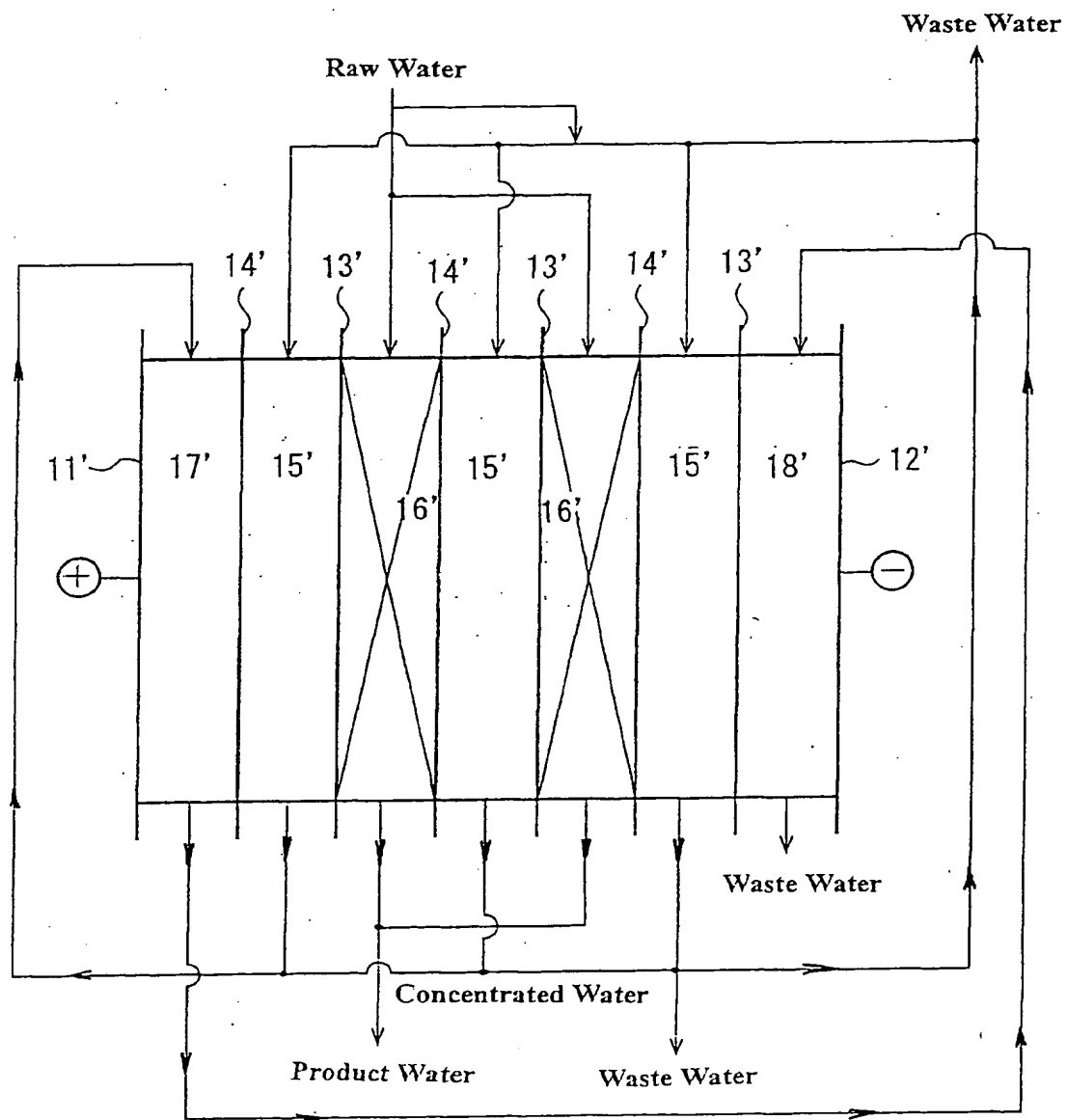
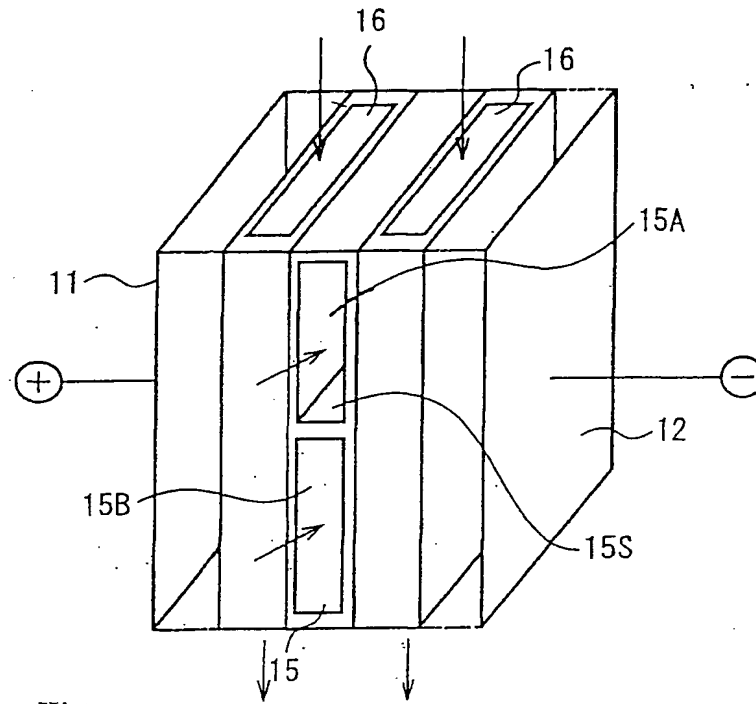


Fig.3a



Raw Water

Fig.3b

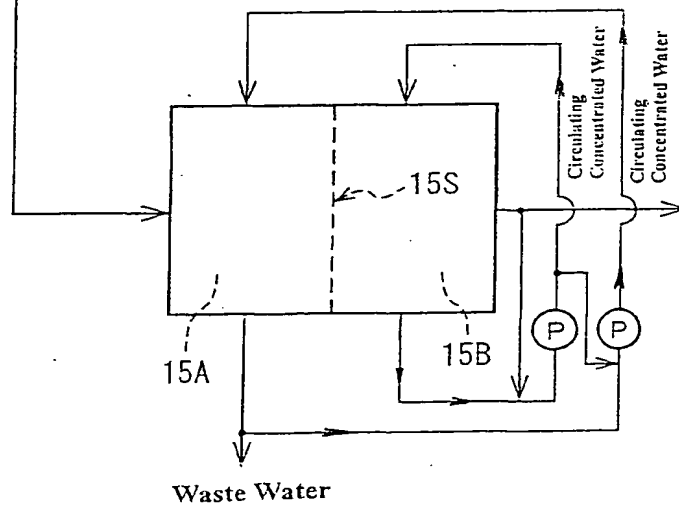
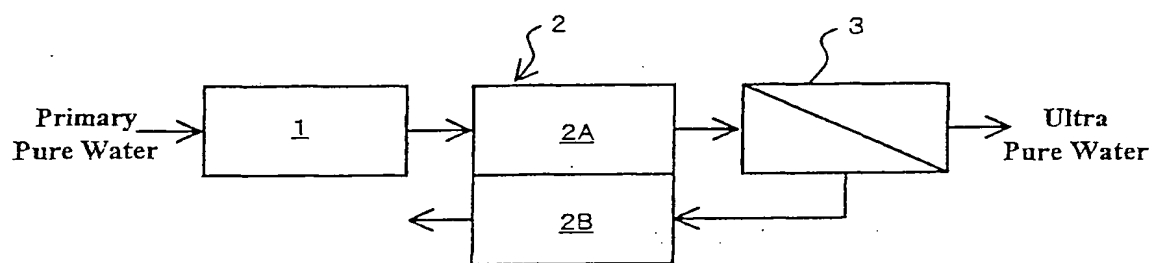


Fig.4





European Patent
Office

EUROPEAN SEARCH REPORT

Application Number
EP 01 31 0901

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.7)
X	US 5 593 563 A (MOULIN JACQUES ET AL) 14 January 1997 (1997-01-14) * column 4, line 1 - line 22; figure 2 * * column 5, line 12 - line 37 * * column 6, line 14 - line 45 *	1,2,4, 10,11, 13,19	B01D61/48 C02F1/00
X	US 3 149 061 A (PARSI EDGARDO J) 15 September 1964 (1964-09-15) * column 2, line 16 - line 41; figure 1 * * column 2, line 61 - column 3, line 10 * * column 4, line 31 - line 64 *	10,11, 13,19	
X	US 5 868 915 A (SPRINGTHORPE PAUL ET AL) 9 February 1999 (1999-02-09) * column 8, line 32 - column 9, line 16 * * column 4, line 37 - line 63; figures 1,2 *	10,11, 13,14,19	
A	WO 97 46492 A (KAKUDA MINORU ;ORGANO CORP (JP); SHINMEI YASUTAKA (JP)) 11 December 1997 (1997-12-11) * abstract; figure 1; examples 1-4 *	1,2,10, 11	TECHNICAL FIELDS SEARCHED (Int.Cl.7) B01D C02F
The present search report has been drawn up for all claims			
Place of search MUNICH		Date of completion of the search 12 April 2002	Examiner Degen, M
<p>CATEGORY OF CITED DOCUMENTS</p> <p>X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document</p> <p>T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application I : document cited for other reasons A : member of the same patent family, corresponding document</p>			

EPO FORM 1503 03 82 (P04001)

**ANNEX TO THE EUROPEAN SEARCH REPORT
ON EUROPEAN PATENT APPLICATION NO.**

EP 01 31 0901

This annex lists the patent family members relating to the patent documents cited in the above-mentioned European search report. The members are as contained in the European Patent Office EDP file on
The European Patent Office is in no way liable for these particulars which are merely given for the purpose of information.

12-04-2002

Patent document cited in search report		Publication date		Patent family member(s)	Publication date
US 5593563	A	14-01-1997	EP	0803474 A2	29-10-1997
			JP	10043554 A	17-02-1998
US 3149061	A	15-09-1964	GB	1013307 A	15-12-1965
			NL	288721 A	
US 5868915	A	09-02-1999	EP	0946301 A1	06-10-1999
			JP	2001500783 T	23-01-2001
			WO	9811987 A1	26-03-1998
			US	5858191 A	12-01-1999
WO 9746492	A	11-12-1997	JP	8150326 A	11-06-1996
			WO	9746492 A1	11-12-1997
			US	6248226 B1	19-06-2001

EPO FORM P0109

For more details about this annex : see Official Journal of the European Patent Office, No. 12/82